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The anion is named by taking

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the elemental name, removing the ending, and adding "ide." For example, F-1 is called fluoride, for the elemental name, fluorine. The "ine" was removed and replaced with "ide." To name a compound, the cation name

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Answer Key and the anion named are added together. For example, NaF is also known as sodium fluoride.

~~Naming Ionic Compounds |
Introduction to Chemistry~~
For binary ionic compounds

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(ionic compounds that contain only two types of elements), the compounds are named by writing the name of the cation first followed by the name of the anion. For example, KCl , an ionic compound that contains K^+

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and Cl^- ions, is named
potassium chloride.

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formulas and naming
compounds can be confusing
because there are different
types of compounds that
follow different rules.
Additionally, some compounds
(H_2O , NH_3 , CH_4 , etc.)
simply have common names

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(Binary Compounds ...~~

We'll learn how to name ionic compounds that have transition metals in them. The names for transition metal compounds often have

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Roman numerals in them,
beca...

~~Naming Ionic Compounds with
Transition Metals
Introduction ...~~

Writing a formula for ionic
compounds containing

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polyatomic ions also involves the same steps as for a binary ionic compound. Write the symbol and charge of the cation followed by the symbol and charge of the anion. Example 5.4. 4: Calcium Nitrate Write the

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~~5.5: Writing Formulas for
Ionic Compounds — Chemistry~~

~~...~~

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Naming Compounds Tutorial
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Ionic Compounds Ternary
Ionic Compounds/Poly-atomic
Ions Naming w/metals that
have more than 1 charge
(Transition Metals)
Molecular Compounds Naming
Acids Metals and Nonmetals
Stairway Of Division on

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Periodic Table C, P, Se, I,
Rn and to the right are non-
metals B ...

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~~Answer Key~~ ~~Learny Kids~~

Here's how to write formulas for binary ionic compounds. We'll see how you have to balance the charges of the two ions so they cancel each other out.

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~~Writing Ionic Formulas:~~

~~Introduction — YouTube~~

Writing Binary Ionic

Compounds. STEP 1: Cation
symbol is written first.

STEP 2: Write the element's
ionic charge (1+, 2+) above
the name. STEP 3: Anion

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Symbol is written next to the positive ion, (single or polyatomic) STEP 4: Write the element's ionic charge (-1 , -2) above the name.

~~3.6 Writing and Naming Ionic
Compounds Flashcards |~~

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Writing a formula for ionic compounds containing polyatomic ions also involves the same steps as for a binary ionic compound. Write the symbol and charge of the cation followed by

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the symbol and charge of the
anion. Example 5.5. 4:

Calcium Nitrate Write the
formula for calcium nitrate.

~~5.5: Writing Formulas for
Ionic Compounds — Chemistry~~

~~...~~

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Answer Key
Ionic compounds consist of a cation and an anion. Naming ionic compounds from their formulas, only the name of the cation first, followed by the anion. These are the rules for naming cations and anions. If communication is

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Answer Key
A type I or Division (Group "A"), the metal, it will be solely responsible, what is meant.

~~Compounds Formula Writing
and Naming ~ Ionic Compounds~~
Rules for naming ionic

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~~Answer Key~~ Always name the metal ion first. Name the nonmetal ion second. Change the ending of the nonmetal or second element by adding -ide.

~~The Rules for Naming Ionic~~

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Answer Key

WRITING AND NAMING IONIC
COMPOUNDS 2 When atoms
combine, its always in
simple whole number ratios
The smallest unit of atomic
combinations that retains
the characteristics of the

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~~Answer Key~~ compound is a molecule 3 The composition of a molecule can be represented in two ways as either an empirical formula or a

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Their name is always written as 2 element names, plus the -ide suffix attached to the second name. Examples of simple binary ionic compounds include potassium oxide and sodium phosphide.

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If the "-ide" suffix doesn't follow a single element name, see instructions for polyatomic ions. For example, "oxide" is a simple oxygen ion, but "hydroxide" and "peroxide" are polyatomic.

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